**Practice Question – Set 1**

**Subject – Chemistry**

**Class - X**

**For question number 1, two statements are given, one labelled Assertion (A) and the other labelled Reason (R). Select the correct answer to these questions from the codes (i), (ii), (iii) and (iv) as given below:**

**(i) Both Assertion (A) and Reason (R) are correct statements, and Reason (R) is the correct explanation of the Assertion (A).**

**(ii) Both Assertion (A) and Reason (R) are correct statements, but Reason (R) is not the correct explanation of the Assertion (A).**

**(iii) Assertion (A) is correct, but Reason (R) is incorrect statement.**

**(iv) Assertion (A) is incorrect, but Reason (R) is correct statement.**

**Q1) Assertion: Silver chloride left open in sunlight turns brown.**

**Reason: The type of reaction involved is decomposition. (1)**

**Q2) Identify the substance oxidised and the substance reduced in each of the following reactions.**

**(i)** $CuO + H_2 \rightarrow Cu + H_2O$

**(ii)** $MnO_2 + 4HCl \rightarrow MnCl_2 + 2H_2O + Cl_2$ **(2)**

**Q3) i) Limestone** $\xrightarrow{\text{heat}}$ **A + B**

**ii)** $A + H_2O \rightarrow$ **Slaked lime**

**a) Identify the exothermic and endothermic reactions from the above.**

**b) Identify B, Write the balanced chemical equation when B reacts with slaked lime.**

**Q4) 2g of ferrous sulphate crystals are heated in a dry boiling tube.**

**(i) List any two observations.**

**(ii) Write the balanced chemical equation for the reaction. (2)**

**Q5) State the type of each of the following reactions stating the reason for your answers:**

**(a)** $Fe_2O_3 + 2Al \rightarrow Al_2O_3 + 2Fe$

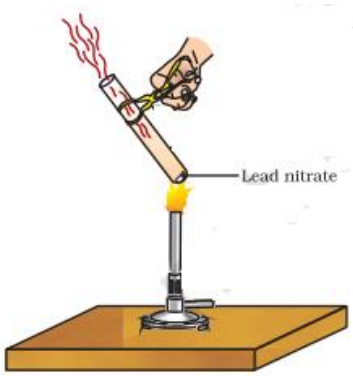
**(b)** $Pb(NO_3)_2 + 2KI \rightarrow PbI_2 + 2KNO_3$ **(2)**

**Q6) (a) A silver article generally turns black when kept in the open for a few days. The article, when rubbed with toothpaste again, starts shining.**

**Why do silver articles turn black when kept in the open for a few days? Define the phenomenon involved.**

**(b) What type of chemical reaction is involved when oily foods get rancid? (3)**

**Q7) The following diagram displays a chemical reaction.**




**(a) How will the colour of the salt change?**
**(b) Name the reddish-brown gas evolved.**
**(c) Write the chemical equation of the reaction that takes place. Mention the state symbols for the reactants and products involved. (3)**