Practice Question – Set 4

Subject – Chemistry

Class - X

For question number 1, two statements are given, one labelled Assertion (A) and the other labelled Reason (R). Select the correct answer to these questions from the codes (i), (ii), (iii) and (iv) as given below:

(i) Both Assertion (A) and Reason (R) are correct statements, and Reason (R) is the correct explanation of the Assertion (A).

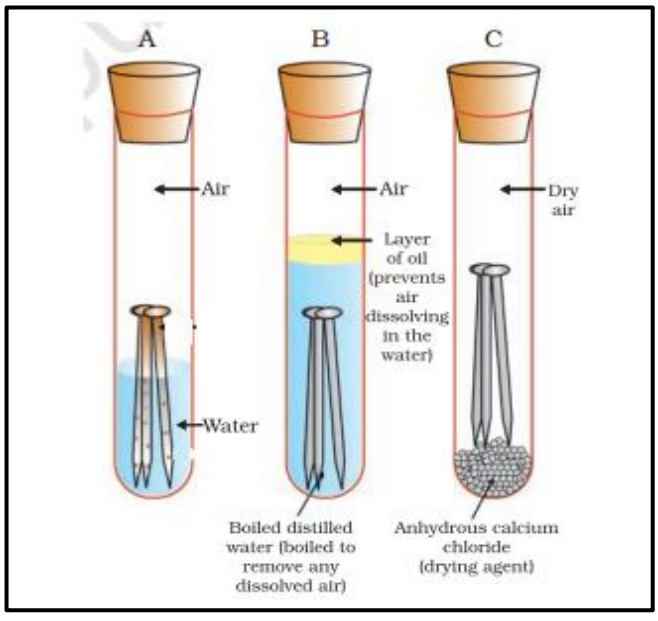
(ii) Both Assertion (A) and Reason (R) are correct statements, but Reason (R) is not the correct explanation of the Assertion (A).

(iii) Assertion (A) is correct, but Reason (R) is incorrect statement.

(iv) Assertion (A) is incorrect, but Reason (R) is correct statement.

Q. 1) **Assertion (A):** Iron displaces zinc from aqueous solution of zinc sulphate .
Reason (R): Zinc is placed above iron in metal activity series. (1)

Q. 2) Observe the following diagram and answer the following questions: (2)



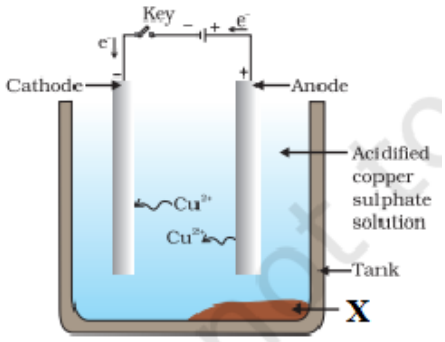
i) What this activity is aimed for?

ii) In which test tube (A/B/C) there will be observable changes and why?

Q. 3) In the following diagram:

i) Mention what anode and cathode are made up of.

ii) What is 'X'? Define. (2)



Q. 4) Name the following: (2)

i) A metal which is best conductor of electricity.

ii) An alloy of copper and tin which is not good conductor of electricity.

Q. 5 i) Write the balanced chemical equation for the following:

The reaction of iron (III) oxide [Fe_2O_3] with heated aluminium which is used to join cracked machine parts.

ii) Draw the electron-dot structure for the formation of magnesium oxide. (Given atomic number: Mg -12, O -16). (2)

Q. 6) i) Name an ore of mercury.

ii) Mention the steps for extraction of mercury from this ore. Also write the balanced chemical equations. (1+2)

Q. 7) Give reason: (3)

(i) Hydrogen gas is not evolved when a metal reacts with nitric acid.

(ii) Metals placed in the top of the activity series are not extracted using the carbon reduction method.

(iii) Aluminium is widely used in making cooking utensils in spite of being a highly reactive metal.